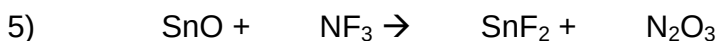
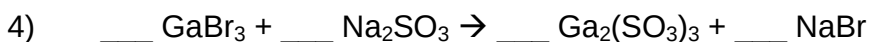
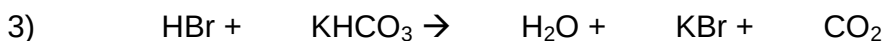
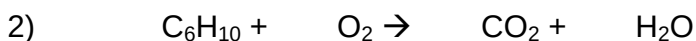


Stoichiometry Practice Worksheet

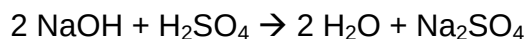
Balancing Equations and Simple Stoichiometry

Balance the following equations:



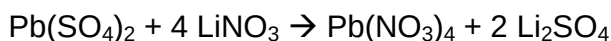
Solve the following stoichiometry grams-grams problems:

- 6) Using the following equation:



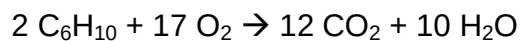
How many grams of sodium sulfate will be formed if you start with 200 grams of sodium hydroxide and you have an excess of sulfuric acid?

- 7) Using the following equation:



How many grams of lithium nitrate will be needed to make 250 grams of lithium sulfate, assuming that you have an adequate amount of lead (IV) sulfate to do the reaction?

Use the following equation to answer questions 8-11:



- 8) If I do this reaction with 35 grams of C_6H_{10} and 45 grams of oxygen, how many grams of carbon dioxide will be formed?
- 9) What is the limiting reagent for problem 6? _____
- 10) How much of the excess reagent is left over after the reaction from problem 6 is finished?
- 11) If 35 grams of carbon dioxide are actually formed from the reaction in problem 6, what is the percent yield of this reaction?

Answer the following stoichiometry-related questions:

- 12) Write the balanced equation for the reaction of acetic acid with aluminum hydroxide to form water and aluminum acetate:
- 13) Using the equation from problem #12, determine the mass of aluminum acetate that can be made if I do this reaction with 125 grams of acetic acid and 275 grams of aluminum hydroxide.
- 14) What is the limiting reagent in problem #13?
- 15) How much of the excess reagent will be left over in problem #13 after the reaction is complete?